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But even the  
humble covalent  
bond - the basic  
building block of



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chemistry - is a consequence of the quantum interactions of electrons. Once these rules were worked out in the 1920s, scientists realised that, for the first time, they had a detailed theory of how chemistry works.

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C<sub>8</sub>H<sub>18</sub> is the formula for octane. The oxygen (O<sub>2</sub>) is consumed from the intake air. Nitrogen (N<sub>2</sub>) is also present in the atmospheric air, but ideally does not participate in

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any reactions (it is quite inert at low temperatures). Note that the combustion products are carbon dioxide ( $\text{CO}_2$ ) and water ( $\text{H}_2\text{O}$ ), if the combustion is 'perfect'. Also, note that there are the same number of each kind ...

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