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Section 55.1.1 Figure 5.1 Different elements can have similar reactions with water. Objectives Compare the wave and particle natures of light. Define a quantum of energy, and explain how it is related to an energy change of matter.

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Models of the Atom. The Rutherford's
model of the atom did not explain how

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an atom can emit light or the chemical properties of an atom.

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Borders. Section 5.1 - Models of the

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Atom. The Rutherford's model of the atom did not explain how an atom can emit light or the chemical properties of an atom. ... Section 5.1 Assessment. Why did Rutherford's atomic model need to be replaced?

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Orbit. TO MOVE AROUND ANOTHER
OBJECT ALONG A PATH.

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Assessment pages 166–169 Section 5.1

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Mastering Concepts 34. Define the following terms. a. frequency Frequency is the number of waves that pass a given point per second. b. wavelength Wavelength is the shortest distance between

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Period _____ MULTIPLE CHOICE. Choose the one alternative that best completes the statement or answers the question.

- 1) The fourth principal energy level has
1) A) 4 orbitals.

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Vocabulary. STUDY. Flashcards. Learn.
Write. Spell. Test. PLAY. Match. Gravity.
Created by. lennm273. Terms in this set
(21) Energy Level. The specific energies
an electron in an atom or other system
can have. Quantum. The amount of
energy needed to move an electron from
one energy level to another.

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Name Class Date Chapter Test A,

continued _____ 8. What is the frequency of a photon whose energy is 3.4×10^{-19} J? ($h = 6.626 \times 10^{-34}$ J·s) a. 8.8×10^{26} Hz b. 5.1×10^{14} Hz c. 1.9×10^{-15} Hz d. 2.3×10^{-52} Hz

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9. When electromagnetic radiation strikes the surface of a metal, electrons are ejected from the metal's ...

Assessment Chapter Test A

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Chapter Test A A. Matching Match each description in Column B with the correct

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term in Column A. Write the letter of the correct description on the line, Column B the number of wave cycles that pass a given point per

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