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GPB 5.24a KEY I. For each of the  
following statements, write "I" for ionic,  
"C" for covalent

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statements, write "I" for ionic, "C" for  
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unit for the following measurements.

Define temperature and mass. a. Length  
b. mass c. volume d. temperature e.

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time f. quantity 9. Perform the following unit conversions: Use the factor label (dimensional analysis) method.

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**UNIT 5 - WELCOME TO CHEMISTRY!**  
AP CHEMISTRY REVIEW WORKSHEET

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(Unit 5 - The Gas Laws) 1. Nitrogen gas has a pressure of 452 mmHg. What is this pressure in atmospheres? In kilopascals? 2. A sample of a certain gas has a volume of 452 mL at 711 mmHg and 26°C. What would be the volume of this same sample of gas if it were measured at STP? 3.

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## **AP CHEMISTRY REVIEW WORKSHEET (Unit 5 The Gas Laws)**

The AP Chemistry course is taught as a continuation of the first year course. As such, units are numbered accordingly. Units 1-8 are covered during the first year course and reviewed over a short time span during the AP course. (Links for these units can be found under the

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"CHEMISTRY" tab above and on the links immediately below this overview.)

## **Units 1-5 Pre-AP Chemistry Review - Mrs. Forest's ...**

Talking concerning Chemistry Unit 5  
Worksheet 1, below we will see several  
related images to give you more ideas.  
chemistry unit 1 worksheet 6

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dimensional analysis, chemistry unit 1 worksheet 3 and chemistry unit 1 worksheet 3 are three main things we want to present to you based on the post title.

## **15 Best Images of Chemistry Unit 5 Worksheet 1 - Chemistry ...**

are representative of each type of

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problem worked in this unit. For more practice on a given type of problem, refer back to Note Taking Guides, worksheets, and quizzes.

1. Calculate the molar mass for ethane,  $C_2H_6$ .  
 $C_2H_6$   
 $2(12.0 \text{ g}) + 6(1.0 \text{ g}) = 30.0 \text{ g}$

2. Convert 45 g Se to moles of Se. ? mol Se  
 $= 45 \text{ g Se} \cdot \frac{1 \text{ mol Se}}{79.0 \text{ g Se}} = 0.57 \text{ mol Se}$

3.

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## **Review Sheet: Unit 7 Name**

### **Vocabulary: Fill in the blanks ...**

CHEMISTRY CP Name: \_\_\_\_\_ KEY \_\_\_\_\_

Period: \_\_\_\_\_ UNIT 5: TEST REVIEW

SHEET MOLE CONVERSIONS,

EMPIRICAL/MOLECULAR FORMULA,

STOICHIOMETRY MOLES AND MOLE

CONVERSIONS 1. Things to Know: a.

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What is a mole? ♦ Avogadro's number of particles; 1 Mole =  $6.02 \times 10^{23}$  particles = gram formula mass of an element/compound b.

**CHEMISTRY CP Name: KEY Period: -  
Oak Park Independent**

Chemistry Unit Review Sheet- Part 2.  
Read the descriptions below and



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determine whether each item describes an endothermic or an exothermic reaction. \_\_\_\_\_ 1. A reaction that produces heat energy \_\_\_\_\_ 2. A reaction that absorbs energy \_\_\_\_\_ 3. Paper burns \_\_\_\_\_ 4

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Unit 5 Review. Show all questions <= ...  
2.5 moles per second; Given the  
following reaction:  $C_3H_8(g) + 5O_2(g) \rightarrow 3CO_2 + 4H_2O(g)$  If the initial  
rate of the reaction produces 2 moles of  
water vapor,  $H_2O$ , per second, what is  
the rate at which propane,  $C_3H_8$ , is  
consumed initially? ?

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## **Unit 5 Review - ScienceGeek.net**

Any other units can be constructed as a combination of fundamental units. For example, velocity could be measured in meters per second (written  $m/s$  or  $m \times s^{-1}$ ), and area could be measured in units of meters squared ( $m^2$ ). When a named unit is defined as a combination of base units, it is called a derived unit. For

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example, the SI unit of energy is the joule (J), which is defined as a  $\text{kg} \times \text{m} \dots$

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