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## **Ideal Gas Law Answer Key**

Ideal Gas Law Worksheet  $PV = nRT$ .

Ideal Gas Law Worksheet  $PV = nRT$ . Use the ideal gas law, " $PV = nRT$ ", and the universal gas constant  $R = 0.0821 \text{ L}\cdot\text{atm}$ .

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to solve the following problems:  $K \cdot \text{mol}$ . If pressure is needed in kPa then convert by multiplying by  $101.3 \text{ kPa} / 1 \text{ atm}$  to get.  $R = 8.31 \text{ kPa} \cdot \text{L} / (\text{K} \cdot \text{mole})$

## **Ideal Gas Law Worksheet $PV = nRT$**

The ideal gas law looks like this:  $PV = nRT$ . The terms in this equation should be mostly familiar to you if you've

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already learned the combined gas law (and the other ones like it). However, if it's not, let's review:  $P$  = the pressure of the gas. In ideal gas equations, this is typically given either in atmospheres or kilopascals.

**The ideal gas law | The Cavalcade o'  
Chemistry**

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Using the Ideal Gas Law ( $PV = nRT$ ), calculate the grams of  $O_2$  produced in the reaction. (Hint: Solve for  $n$  and then convert moles to grams. Don't forget to convert your temperature from Celsius...

**Ideal Gas Law Questions and  
Answers | Study.com**  
Ideal Gas Law Practice Worksheet

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Solutions to the Ideal gas law practice worksheet: The ideal gas law states that  $PV=nRT$ , where  $P$  is the pressure of a gas,  $V$  is the volume of the gas,  $n$  is the number of moles of gas present,  $R$  is the ideal gas constant, and  $T$  is the temperature of the gas in Kelvins.

## **Ideal Gas Law Practice Worksheet**



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## **Answer Key**

THE IDEAL GAS LAW and DENSITY.  $PV = nRT$  where. pressure in atmosphere.  
volume in liters. = number of moles of  
gas. Universal Gas Constant = 0.0821  
L.atm/mol.K. Kelvin temperature.

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Ideal Gas Law Worksheet  $PV = nRT$ . Use the ideal gas law, and the universal gas

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constant to solve the following  
problems: with atm:  $R = 0.0821 \text{ L}\cdot\text{atm} / (\text{K}\cdot\text{mol})$  with kPa:  $R = 8.31 \text{ L}\cdot\text{kPa} / (\text{K}\cdot\text{mole})$  1) If I have 4 moles of a gas at a pressure of 5.6 atm and a volume of 12 liters, what is the temperature?

**Ideal Gas Law Worksheet  $PV = nRT$**   
Ideal Gas Law and Stoichiometry Use the

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following reaction to answer the next few questions:  $2 \text{C}_8\text{H}_{18}(\text{l}) + 25 \text{O}_2(\text{g}) \dots > 16 \text{CO}_2(\text{g}) + 18 \text{H}_2\text{O}(\text{g})$  The above reaction is the reaction between gasoline (octane) and oxygen that occurs inside automobile engines.

**Gas Laws STUDY GUIDE Due:  
February 12th**

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The Ideal Gas Law can be re-arranged to calculate the molar mass of unknown gases.  $PV = nRT$   $n = \frac{\text{mass (g)}}{\text{molar mass (g/mol)}}$   $PV = \frac{\text{mass}}{\text{molar mass}} (RT)$   $\text{mass} \times R \times T = \text{molar mass} \times P \times V$  Knowing that the units for density are mass/volume, re-write this equation so that it equates density with molar mass.

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## **Worksheet 7 - Ideal Gas Law I. Ideal Gas Law Ideal Gas Law ...**

$P_1V_1/T_1 = P_2V_2/T_2$ . Ideal Gas Law: An ideal gas must follow the Kinetic Molecular Theory of Gases. We have talked about four variables that affect the behavior of gases. The four gas variables are:...

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Ideal Gas Law With the combined gas

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law, you can solve problems with three variables: pressure, volume, and temperature. The combined gas law assumes that the amount of gas does not vary. You cannot use the combined gas law to calculate the number of moles of a gas in a fixed volume at a known temperature and pressure.

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## **14.3 Ideal Gases - Henry County School District**

IDEAL GAS LAW Use the ideal Gas Law below to solve the following problems.  
pressure in atmospheres volume in liters  
number of moles L atm Universal Gas  
Constant = 0.0821 mol. K Kelvin  
temperature 1. 2. 3. 4.

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